Objectives

After completing this section, you should be able to

1. identify the region of the infrared spectrum in which the carbonyl absorption of aldehydes and ketones is found.
2. identify the region of the infrared spectrum in which the two characteristic C=H absorptions of aldehydes are found.
3. use a table of characteristic absorption frequencies to assist in the determination of the structure of an unknown aldehyde or ketone, given its infrared spectrum and other spectral or experimental data.
4. identify the region of a proton NMR spectrum in which absorptions caused by the presence of aldehydic protons and protons attached to the α-carbon atoms of aldehydes and ketones occur.
5. identify two important fragmentations that occur when aliphatic aldehydes and ketones are subjected to analysis by mass spectrometry.

Key Terms

Make certain that you can define, and use in context, the key term below.

- McLafferty rearrangement

Study Notes

The appearance of a strong absorption at 1660–1770 cm$^{-1}$ in the infrared spectrum of a compound is a clear indication of the presence of a carbonyl group. Although you need not remember the detailed absorptions it is important that you realize that the precise wavenumber of the infrared absorption can often provide some quite specific information about the environment of the carbonyl group in a compound. Notice how conjugation between a carbonyl group and a double bond (α, β-unsaturated aldehyde or ketone or aromatic ring) lowers the absorption by about 25–30 cm$^{-1}$.

You may wish to review the McLafferty rearrangement and the alpha cleavage in Section 12.3.

IR Spectra

The carbonyl stretching vibration band C=O of saturated aliphatic ketones appears:

**C=O stretch**

- aliphatic ketones 1715 cm$^{-1}$
- alpha, beta-unsaturated ketones 1685-1666 cm$^{-1}$

Figure 8. shows the spectrum of 2-butanone. This is a saturated ketone, and the C=O band appears at 1715.
If a compound is suspected to be an aldehyde, a peak always appears around 2720 cm\(^{-1}\) which often appears as a shoulder-type peak just to the right of the alkyl C–H stretches.

**H–C=O stretch** 2830-2695 cm\(^{-1}\)

**C=O stretch**
- aliphatic aldehydes 1740-1720 cm\(^{-1}\)
- alpha, beta-unsaturated aldehydes 1710-1685 cm\(^{-1}\)

Figure 9. shows the spectrum of butyraldehyde.

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**NMR Spectra**

Hydrogens attached to carbon adjacent to the \(sp^2\) hybridized carbon in aldehydes and ketones usually show up 2.0-2.5 ppm.
Aldehyde hydrogens are highly shielded and appear far downfield as 9-10 ppm.

Chemical shift of each protons is predicted by $^1H$ chemical shift ranges ($H_a$): chemical shift of methyl groups (1.1 ppm). ($H_b$) The chemical shift of the -CH- group move downfield due to effect an adjacent aldehyde group: (2.4 ppm). The chemical shift of aldehyde hydrogen is highly shielded (9.6 ppm).

4) Splitting pattern is determined by (N+1) rule: $H_a$ is split into two peaks by $H_b$ (#of proton=1). $H_b$ has the septet pattern by $H_a$ (#of proton=6). $H_c$ has one peak.(Note that $H_c$ has doublet pattern by $H_b$ due to vicinal proton-proton coupling.)

Mass Spectra

Aldehydes and ketones generally give moderately intense signals due to their molecular ions, $\text{M}^+$. Thus the determination of the molecular weight of a ketone by mass spectroscopy usually is not difficult. Furthermore, there are some characteristic fragmentation patterns that aid in structural identification. These are:

- $\alpha$ cleavage

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\begin{align*}
\text{R} & \quad \text{C} = \hat{\text{O}}^+ \\
\text{R} & \quad \text{R} - \text{C} = \hat{\text{O}}^+ + \text{R}^-
\end{align*}
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- Transfer of $\gamma$ hydrogen with $\beta$ cleavage (McLafferty rearrangement)
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