The nitrogen rule states that a molecule that has no or even number of nitrogen atoms has an even nominal mass,
whereas a molecule that has an odd number of nitrogen atoms has an odd nominal mass.

eg. 1:

```
H
H—C—O—H
H
```

molecular formula = CH₄O
nominal mass = (1x12) + (4x1) + (1x16)
= 32

# N atoms = 0
nominal mass = 32 (even #)

eg. 2:

```
H
H—C—N
H
```

molecular formula = CH₅N
nominal mass = (1x12) + (5x1) + (1x14)
= 31

# N atoms = 1 (odd #)
nominal mass = 31 (odd #)

eg. 3:
molecular formula = C₂H₆N₂
nominal mass = (2x12) + (6x1) + (2x14)
= 58

# N atoms = 2 (even #)
nominal mass = 58 (even #)

Contributors

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