Electron-rich, the hydride ion (1) should in theory be a base and a nucleophile.

\[
\begin{align*}
\text{H}^- \\
\text{1}
\end{align*}
\]

In practice, however, the hydride ion is strongly basic but not nucleophilic. A reagent that, in a reaction, acts as the hydride ion would if it were nucleophilic is called a hydride reagent or hydride equivalent. The most common hydride reagents are the reducing agents sodium borohydride (NaBH₄) and lithium aluminumhydride (LiAlH₄).

eg:

Contributors

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