This page looks at how fragmentation patterns are formed when organic molecules are fed into a mass spectrometer, and how you can get information from the mass spectrum.

**The formation of molecular ions**

When the vaporized organic sample passes into the ionization chamber of a mass spectrometer, it is bombarded by a stream of electrons. These electrons have a high enough energy to knock an electron off an organic molecule to form a positive ion. This ion is called the molecular ion - or sometimes the parent ion. The molecular ion is often given the symbol $\ce{M^{+}}$ or $\ce{M^\cdot +}$- the dot in this second version represents the fact that somewhere in the ion there will be a single unpaired electron. That's one half of what was originally a pair of electrons - the other half is the electron which was removed.

**Fragmentation**

The molecular ions are energetically unstable, and some of them will break up into smaller pieces. The simplest case is that a molecular ion breaks into two parts - one of which is another positive ion, and the other is an uncharged free radical.

$$\ce{M^\cdot + \rightarrow X^+ + Y^\cdot}$$

The uncharged free radical won't produce a line on the mass spectrum. Only charged particles will be accelerated, deflected and detected by the mass spectrometer. These uncharged particles will simply get lost in the machine - eventually, they get removed by the vacuum pump. The ion, $\ce{X^+}$, will travel through the mass spectrometer just like any other positive ion - and will produce a line on the stick diagram. All sorts of fragmentations of the original molecular ion are possible - and that means that you will get a whole host of lines in the mass spectrum. For example, the mass spectrum of pentane looks like this:

It's important to realize that the pattern of lines in the mass spectrum of an organic compound tells you something quite different from the pattern of lines in the mass spectrum of an element. With an element, each line represents a different isotope of that element. With a compound, each line represents a different fragment produced when the molecular ion breaks up.
In the stick diagram showing the mass spectrum of pentane, the line produced by the heaviest ion passing through the machine (at m/z = 72) is due to the molecular ion. The tallest line in the stick diagram (in this case at m/z = 43) is called the **base peak**. This is usually given an arbitrary height of 100, and the height of everything else is measured relative to this. The base peak is the tallest peak because it represents the commonest fragment ion to be formed - either because there are several ways in which it could be produced during fragmentation of the parent ion, or because it is a particularly stable ion.

### Using fragmentation patterns

This section will ignore the information you can get from the molecular ion (or ions). That is covered in three other pages which you can get at via the mass spectrometry menu. You will find a link at the bottom of the page.

**Example (PageIndex(1)): Mass Spectrum of Pentane**

Let's have another look at the mass spectrum for pentane:

What causes the line at m/z = 57?

**Solution**

How many carbon atoms are there in this ion? There can't be 5 because 5 x 12 = 60. What about 4? 4 x 12 = 48. That leaves 9 to make up a total of 57. How about \( \text{C}_4\text{H}_9^+ \) then?

\( \text{C}_4\text{H}_9^+ \) would be \([\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2]^+\), and this would be produced by the following fragmentation:

\[
[\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2]^+ \rightarrow [\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2]^+ + \cdot\text{CH}_3
\]

The methyl radical produced will simply get lost in the machine.

The line at m/z = 43 can be worked out similarly. If you play around with the numbers, you will find that this corresponds to a break producing a 3-carbon ion:

\[
[\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2]^+ \rightarrow [\text{CH}_3\text{CH}_2\text{CH}_2]^+ + \cdot\text{CH}_2\text{CH}_3
\]
The line at $m/z = 29$ is typical of an ethyl ion, $[\text{CH}_3\text{CH}_2]^+$:

$$[\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3]^+ \rightarrow [\text{CH}_3\text{CH}_2]^+ + \text{CH}_2\text{CH}_3$$

The other lines in the mass spectrum are more difficult to explain. For example, lines with $m/z$ values 1 or 2 less than one of the easy lines are often due to loss of one or more hydrogen atoms during the fragmentation process.

Example \(\PageIndex{2}\): Pentan-3-one

This time the base peak (the tallest peak - and so the commonest fragment ion) is at $m/z = 57$. But this isn't produced by the same ion as the same $m/z$ value peak in pentane.

If you remember, the $m/z = 57$ peak in pentane was produced by $[\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2]^+$. If you look at the structure of pentan-3-one, it's impossible to get that particular fragment from it.

Work along the molecule mentally chopping bits off until you come up with something that adds up to 57. With a small amount of patience, you'll eventually find $[\text{CH}_3\text{CH}_2\text{CO}]^+$ - which is produced by this fragmentation:

$$[\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3]^+ \rightarrow [\text{CH}_3\text{CH}_2\text{CO}]^+ + \text{CH}_2\text{CH}_3$$

You would get exactly the same products whichever side of the CO group you split the molecular ion. The $m/z = 29$ peak is produced by the ethyl ion - which once again could be formed by splitting the molecular ion either side of the CO group.

Peak heights and the stability of ions

The more stable an ion is, the more likely it is to form. The more of a particular sort of ion that's formed, the higher its peak height will be. We'll look at two common examples of this.

Examples involving carbocations (carbonium ions)

Summarizing the most important conclusion from the page on carbocations:
Order of stability of carbocations

primary < secondary < tertiary

Applying the logic of this to fragmentation patterns, it means that a split which produces a secondary carbocation is going to be more successful than one producing a primary one. A split producing a tertiary carbocation will be more successful still. Let's look at the mass spectrum of 2-methylbutane. 2-methylbutane is an isomer of pentane - isomers are molecules with the same molecular formula, but a different spatial arrangement of the atoms.

Look first at the very strong peak at m/z = 43. This is caused by a different ion than the corresponding peak in the pentane mass spectrum. This peak in 2-methylbutane is caused by:

\[
\begin{align*}
\left[\text{CH}_3\text{C}^+\text{CH}_2\text{CH}_3\right] & \rightarrow \text{CH}_3\text{C}^+\text{H} + \cdot\text{CH}_2\text{CH}_3 \\
\end{align*}
\]

The ion formed is a secondary carbocation - it has two alkyl groups attached to the carbon with the positive charge. As such, it is relatively stable. The peak at m/z = 57 is much taller than the corresponding line in pentane. Again a secondary carbocation is formed - this time, by:

\[
\begin{align*}
\left[\text{CH}_3\text{C}^+\text{C}=\cdot\text{CH}_3\right] & \rightarrow \text{CH}_3\text{C}^+\text{CH}_2\text{CH}_3 + \cdot\text{CH}_3 \\
\end{align*}
\]

You would get the same ion, of course, if the left-hand CH$_3$ group broke off instead of the bottom one as we've drawn it. In these two spectra, this is probably the most dramatic example of the extra stability of a secondary carbocation.

Examples involving acylium ions, [RCO]$^+$

Ions with the positive charge on the carbon of a carbonyl group, C=O, are also relatively stable. This is fairly clearly seen in the mass spectra of ketones like pentan-3-one.
The base peak, at m/z=57, is due to the \([\text{CH}_3\text{CH}_2\text{CO}]^+\) ion. We've already discussed the fragmentation that produces this.

### Using mass spectra to distinguish between compounds

Suppose you had to suggest a way of distinguishing between pentan-2-one and pentan-3-one using their mass spectra.

- **pentan-2-one** \(\text{CH}_3\text{COCH}_2\text{CH}_2\text{CH}_3\)
- **pentan-3-one** \(\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3\)

Each of these is likely to split to produce ions with a positive charge on the CO group. In the pentan-2-one case, there are two different ions like this:

- \([\text{CH}_3\text{CO}]^+\)
- \([\text{COCH}_2\text{CH}_2\text{CH}_3]^+\)

That would give you strong lines at m/z = 43 and 71.

With pentan-3-one, you would only get one ion of this kind:

- \([\text{CH}_3\text{CH}_2\text{CO}]^+\)

In that case, you would get a strong line at 57. You don't need to worry about the other lines in the spectra - the 43, 57 and 71 lines give you plenty of difference between the two. The 43 and 71 lines are missing from the pentan-3-one spectrum, and the 57 line is missing from the pentan-2-one one.

The two mass spectra look like this:
Computer matching of mass spectra

As you've seen, the mass spectrum of even very similar organic compounds will be quite different because of the different fragmentations that can occur. Provided you have a computer data base of mass spectra, any unknown spectrum can be computer analysed and simply matched against the data base.

Contributors

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