Objectives

After completing this section, you should be able to

1. discuss the wide occurrence of esters in nature, and their important commercial uses, giving one example of an ester linkage in nature, and one example of a commercially important ester.
2. write an equation to describe the hydrolysis of an ester under acidic or basic conditions.
3. identify the products formed from the hydrolysis of an given ester.
4. identify the reagents that can be used to bring about ester hydrolysis.
5. identify the structure of an unknown ester, given the products of its hydrolysis.
6. write the mechanism of alkaline ester hydrolysis.
7. write the mechanism of acidic ester hydrolysis.
8. write an equation to describe the reduction of an ester with lithium aluminum hydride.
9. identify the product formed from the reduction of a given ester (or lactone) with lithium aluminum hydride.
10. identify the ester, the reagents, or both, that should be used to prepare a given primary alcohol.
11. write a detailed mechanism for the reduction of an ester by lithium aluminum hydride.
12. identify diisobutylaluminum hydride as a reagent for reducing an ester to an aldehyde, and write an equation for such a reaction.
13. write an equation to describe the reaction of an ester with a Grignard reagent.
14. identify the product formed from the reaction of a given ester with a given Grignard reagent.
15. identify the ester, the Grignard reagent, or both, needed to prepare a given tertiary alcohol.
16. write a detailed mechanism for the reaction of an ester with a Grignard reagent.

Key Terms

Make certain that you can define, and use in context, the key terms below.

- lactone
- saponification

Study Notes

Many esters have characteristic aromas and flavours. Some examples are listed below.

**Basic structure:**

\[ R\text{O-}\text{O-R'} \]
Table 21.1 Structures and characteristic odours of selected esters

<table>
<thead>
<tr>
<th>IUPAC name</th>
<th>R</th>
<th>R'</th>
<th>Aroma</th>
</tr>
</thead>
<tbody>
<tr>
<td>octyl ethanoate</td>
<td>CH₃</td>
<td>CH₃(CH₂)₆CH₂</td>
<td>orange</td>
</tr>
<tr>
<td>propyl ethanoate</td>
<td>CH₃</td>
<td>CH₃CH₂CH₂</td>
<td>pear</td>
</tr>
<tr>
<td>2-methylpropyl propanoate</td>
<td>CH₃CH₂</td>
<td>(CH₃)₂CHCH₂</td>
<td>rum</td>
</tr>
<tr>
<td>methyl butanoate</td>
<td>CH₃CH₂CH₂</td>
<td>CH₃</td>
<td>apple</td>
</tr>
<tr>
<td>ethyl butanoate</td>
<td>CH₃CH₂CH₂</td>
<td>CH₃CH₂</td>
<td>pineapple</td>
</tr>
</tbody>
</table>

A "lactone" is a cyclic ester and has the general structure

![Lactone structure](image)

By recognizing that the steps in the acidic hydrolysis of an ester are exactly the same as those in a Fischer esterification (but in the reverse order!), you can again minimize the amount of memorization that you must undertake. The details of both mechanisms can be deduced from the knowledge that both reactions are acid-catalyzed nucleophilic acyl substitutions.

**Introduction**

Esters are readily synthesized and naturally abundant contributing to the flavors and aromas in many fruits and flowers.

They also make up the bulk of animal fats and vegetable oils—glycerides (fatty acid esters of glycerol). Soap is produced by a saponification (basic hydrolysis) reaction of a fat or oil.
Esters are also present in a number of important biological molecules and have several commercial and synthetic application. For example, polyester molecules make excellent fibers and are used in many fabrics. A knitted polyester tube, which is biologically inert, can be used in surgery to repair or replace diseased sections of blood vessels. PET is used to make bottles for soda pop and other beverages. It is also formed into films called Mylar. When magnetically coated, Mylar tape is used in audio- and videocassettes. Synthetic arteries can be made from PET, polytetrafluoroethylene, and other polymers.

The most important polyester, polyethylene terephthalate (PET), is made from terephthalic acid and ethylene glycol monomers:

\[
\begin{align*}
  n \text{ HOCH}_2\text{CH}_2\text{OH} & \quad + \quad n \text{ HOOC} & \quad \longrightarrow \\
  & \quad \text{Terephthalic acid} & \quad \text{Ethylene glycol}
\end{align*}
\]

\[
\begin{align*}
  \left[ \text{CO} & \quad \text{COOCH}_2\text{CH}_2\text{O} \right]_n \\
  & \quad n \text{ H}_2\text{O} & \quad + \quad n \text{ H}_2\text{O}
\end{align*}
\]

Polyethylene terephthalate

Preparation of Esters

Carboxylic acids can react with alcohols to form esters
Acid chlorides react with alcohols to form esters

\[
\text{Acid Chloride} + \text{Alcohol} \rightarrow \text{Ester} + \text{HCl}
\]

Acid Anhydrides react with alcohols to form esters

\[
\text{Acid Anhydride} + \text{Alcohol} \rightarrow \text{Ester} + \text{Carboxylic acid}
\]

**Conversion of Ester into Carboxylic acids: Hydrolysis**

Esters can be cleaved back into a carboxylic acid and an alcohol by reaction with water and a catalytic amount of acid.

\[
\text{Ester} + \text{H}_2\text{O} + \text{H}_3\text{O}^+ \rightarrow \text{Carboxylic Acid} + \text{Alcohol}
\]

**Example 21.6.1**

\[
\text{Ester} \quad \text{H}_2\text{O} \quad \text{H}_3\text{O}^+ \rightarrow \text{Carboxylic Acid} + \text{Alcohol}
\]

**Mechanism**

1) Protonation of the Carbonyl

\[
\text{Ester} \quad \text{H}_2\text{O} \quad \text{H}_3\text{O}^+ \rightarrow \text{Carboxylic Acid} + \text{Alcohol}
\]

2) Nucleophilic attack by water
3) Proton transfer

4) Leaving group removal

**Esters** can be cleaved back into a carboxylic acid and an alcohol by reaction with water and a base. The reaction is called a saponification from the Latin *sapo* which means soap. The name comes from the fact that soap used to made by the ester hydrolysis of fats. Due to the basic conditions a carboxylate ion is made rather than a **carboxylic acid**.

Example 21.6.2

**Mechanism**

1) Nucleophilic attack by hydroxide
2) Leaving group removal

3) Deprotonation

**Conversion of Esters into Amides: Aminolysis**

Esters reaction with ammonia and alkyl amines to yield amides.

\[
\text{R-C}O\text{-CH}_3 + \text{NH}_3 \rightarrow \text{R-C}N\text{H}_2 + \text{CH}_3\text{OH}
\]

**Conversion of Ester into Alcohols: Reduction**

Esters can be converted to 1° alcohols using LiAlH₄, while sodium borohydride (\(\text{NaBH}_4\)) is not a strong enough reducing agent to perform this reaction.

Example 21.6.3
Esters can be converted to aldehydes using diisobutylaluminum hydride (DIBAH)

The reaction is usually carried out at \(-78^\circ C\) to prevent reaction with the aldehyde product.
Example 21.6.4

Addition of Grignard reagents convert esters to $3^\circ$ alcohols.

In effect the Grignard reagent adds twice.

Example 21.6.5

Mechanism

1) Nucleophilic attack

2) Leaving group removal
3) Nucleophilic attack

\[
\begin{align*}
\text{R''} & \text{MgX} \\
\text{R} & \text{R''} \\
\end{align*}
\]

4) Protonation

\[
\begin{align*}
\text{R} & \text{R''} \\
\text{R} & \text{R''} + \text{H}_2\text{O} \\
\end{align*}
\]

Exercises

Questions

Q21.6.1

Why is the alkaline hydrolysis of an ester not a reversible process? Why doesn't the reaction with a hydroxide ion and a carboxylic acid produce an ester?

Q21.6.2

Draw the product of the reaction between the following molecule and LiAlH₄, and the product of the reaction between the following molecule and DIBAL.

Q21.6.3

How might you Prepare the following molecules from esters and Grignards?

(a)
Solutions

S21.6.1

The reaction between a carboxylic acid and a hydroxide ion is an acid base reaction, which produces water and a carboxylate anion.

S21.6.2
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