Innovative study for the prediction of bond order in case of oxide based acid radicals have been discussed here\textsuperscript{1-3}.

**In case of oxide based acid radicals**

**Bond Order (B.O.)** = Valency of the peripheral atom + (Charge on Acid Radical / Total number of peripheral atoms)

Eg.:

- $\text{ClO}_4^-$: (Valency of one Peripheral atom Oxygen = 2, Charge on acid radical = -1, Total Number of Peripheral atoms = 04), Therefore B.O. = 2 + (-1/4) = 1.75
- $\text{ClO}_3^-$: (Valency of one Peripheral atom Oxygen = 2, Charge on acid radical = -1, Total Number of Peripheral atoms = 03), Therefore B.O. = 2 + (-1/3) = 1.66
- $\text{ClO}_2^-$: (Valency of one Peripheral atom Oxygen = 2, Charge on acid radical = -1, Total Number of Peripheral atoms = 02), Therefore B.O. = 2 + (-1/2) = 1.5
- $\text{AsO}_4^{3-}$: (Valency of one Peripheral atom Oxygen = 2, Charge on acid radical = -3, Total Number of Peripheral atoms = 04), Therefore B.O. = 2 + (-3/4) = 1.25
- $\text{AsO}_3^{3-}$: (Valency of one Peripheral atom Oxygen = 2, Charge on acid radical = -3, Total Number of Peripheral atoms = 03), Therefore B.O. = 2 + (-3/3) = 1.0
- $\text{SO}_4^{2-}$: (Valency of Peripheral atom Oxygen = 2, Charge on acid radical = -2, Number of Peripheral atoms = 04), Therefore B.O. = 2 + (-2/4) = 1.5
- $\text{SO}_3^{2-}$: (Valency of Peripheral atom Oxygen = 2, Charge on acid radical = -2, Number of Peripheral atoms = 03), Therefore B.O. = 2 + (-2/3) = 1.33
- $\text{PO}_4^{3-}$: (Valency of Peripheral atom Oxygen = 2, Charge on acid radical = -3, Number of Peripheral atoms = 04), Therefore B.O. = 2 + (-3/4) = 1.25
- $\text{BO}_3^{3-}$: (Valency of Peripheral atom Oxygen = 2, Charge on acid radical = -3, Number of Peripheral atoms = 03), Therefore B.O. = 2 + (-3/3) = 1
- $\text{CO}_3^{2-}$: (Valency of Peripheral atom Oxygen = 2, Charge on acid radical = -2, Number of Peripheral atoms = 03), Therefore B.O. = 2 + (-2/3) = 1.33
- $\text{SiO}_4^{4-}$: (Valency of Peripheral atom Oxygen = 2, Charge on acid radical = -4, Number of Peripheral atoms = 04), Therefore B.O. = 2 + (-4/4) = 1

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**Relation (Bond order vs. Bond length, Bond Strength, Bond energy, Thermal stability and Reactivity)**

B.O. $\propto \frac{1}{\text{Bond length or Bond distance}}$;

B.O. $\propto \text{Bond strength}$;

B.O. $\propto \text{Bond Energy}$;

B.O. $\propto \text{Thermal Stability}$; B.O. $\propto \frac{1}{\text{ Reactivity}}$
Correlation (Literature values of bond-distances of some oxide based acid radicals with their predicted bond order values)

Literature values of the Cl-O average bond lengths in ClO_4^- , ClO_3^- and ClO_2^- ; As-O bond lengths in AsO_4^{3-} and AsO_3^{3-} with respect to their bond order values suggest that with increasing bond-order M-O bond length (Where M = Cl, As etc.) decreases which is shown in Table-1.

Table 1: Bond-distances and their predicted bond order values

<table>
<thead>
<tr>
<th>Oxide Based Acid Radicals</th>
<th>Bond-Order Values</th>
<th>Avg. M-O Bond-Distances As per Literature (Å)</th>
<th>Remarks</th>
</tr>
</thead>
<tbody>
<tr>
<td>ClO_4^-</td>
<td>1.75</td>
<td>1.50</td>
<td></td>
</tr>
<tr>
<td>ClO_3^-</td>
<td>1.6</td>
<td>1.57</td>
<td></td>
</tr>
<tr>
<td>ClO_2^-</td>
<td>1.5</td>
<td>1.64</td>
<td>Increasing Bond-Order decreases Bond Length</td>
</tr>
<tr>
<td>AsO_4^{3-}</td>
<td>1.25</td>
<td>1.75</td>
<td></td>
</tr>
<tr>
<td>AsO_3^{3-}</td>
<td>1.0</td>
<td>1.77</td>
<td></td>
</tr>
</tbody>
</table>

References

1. Arijit Das, ‘New Methods for Prediction of Bond Order of Mono and Diatomic Homo and Hetero Nuclear Molecules or Ions Having (1-20)e-s and Oxide Based Acid Radicals Without MOT – a Rapid Innovative Approach’, IJAR, 2013, 3(11), 41-43, ISSN-2249-555X.

External Links

- [https://communities.wbr/acs.org/docs/DOC-46667](https://communities.wbr/acs.org/docs/DOC-46667)
- [https://communities.acs.org/docs/DOC-45853](https://communities.acs.org/docs/DOC-45853)
Contributor

Dr. Arijit Das, Ph.D. (Inorganic Chemistry), MACS (Invited, USA), SFICS, MISC, MIAFS (India), Assistant Professor, Department of Chemistry, Ramthakur College, Agartala, Tripura (W), Tripura, India, Pin-799003.