Following are examples of compounds listed by functional group, which demonstrate patterns which can be seen in mass spectra of compounds ionized by electron impact ionization. These examples do not provide information about the fragmentation mechanisms that cause these patterns. Additional information can be found in mass spectrometry reference books.

**Alcohol**

An alcohol's molecular ion is small or non-existent. Cleavage of the C-C bond next to the oxygen usually occurs. A loss of \( \text{H}_2\text{O} \) may occur as in the spectra below.

3-Pentanol (C\(_5\)H\(_{12}\)O) with MW = 88.15

\[ \text{CH}_3\text{CH}_2\text{C}-\text{CH}_2\text{CH}_3 \]

\[ \text{OH} \]

\[ -\text{H}_2\text{O} \]

\[ \text{m/z} = 88 \]

**Aldehyde**

Cleavage of bonds next to the carboxyl group results in the loss of hydrogen (molecular ion less 1) or the loss of CHO (molecular ion less 29).

3-Phenyl-2-propenal (C\(_9\)H\(_8\)O) with MW = 132.16

\[ \text{m/z} = 132 \]
**Alkane**

Molecular ion peaks are present, possibly with low intensity. The fragmentation pattern contains clusters of peaks 14 mass units apart (which represent loss of (CH2)nCH3).

Hexane (C6H14) with MW = 86.18

![Graph of Alkane](image)

**Amide**

Primary amides show a base peak due to the McLafferty rearrangement.

3-Methylbutyramide (C5H11NO) with MW = 101.15

![Graph of Amide](image)

**Amine**

Molecular ion peak is an odd number. Alpha-cleavage dominates aliphatic amines.

n-Butylamine (C4H11N) with MW = 73.13
Another example is a secondary amine shown below. Again, the molecular ion peak is an odd number. The base peak is from the C-C cleavage adjacent to the C-N bond.

n-Methylbenzylamine (C₈H₁₁N) with MW = 121.18

Aromatic

Molecular ion peaks are strong due to the stable structure.

Naphthalene (C₁₀H₈) with MW = 128.17
Carboxylic Acid

In short chain acids, peaks due to the loss of OH (molecular ion less 17) and COOH (molecular ion less 45) are prominent due to cleavage of bonds next to C=O.

2-Butenoic acid (C₄H₆O₂) with MW = 86.09

Ester

Fragments appear due to bond cleavage next to C=O (alkoxy group loss, -OR) and hydrogen rearrangements.

Ethyl acetate (C₄H₈O₂) with MW = 88.11
**Ether**

Fragmentation tends to occur alpha to the oxygen atom (C-C bond next to the oxygen).

Ethyl methyl ether ($\text{C}_3\text{H}_8\text{O}$) with MW = 60.10

**Halide**

The presence of chlorine or bromine atoms is usually recognizable from isotopic peaks.

1-Bromopropane ($\text{C}_3\text{H}_7\text{Br}$) with MW = 123.00
**Ketone**

Major fragmentation peaks result from cleavage of the C-C bonds adjacent to the carbonyl.

4-Heptanone \((C_7H_{14}O)\) with MW = 114.19

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