Halogenation is an example of electrophilic aromatic substitution. In electrophilic aromatic substitutions, a benzene is attacked by an electrophile which results in substitution of hydrogens. However, halogens are not electrophilic enough to break the aromaticity of benzenes, which require a catalyst to activate.

**Activation of Halogen**

(where X = Br or Cl, we will discuss later why other members of the halogen family fluorine and iodine are not used in halogenation of benzenes)

Halogens are not reactive enough on their own to react with an aromatic ring. Halogens need a Lewis Acidic catalyst to activate them to become a very strong electrophile. Examples of these activated halogens are ferric halides (FeX₃) and aluminum halides (AlX₃) where X = Br or Cl.

In order to make bromine electrophilic enough to react with benzene, we use the aid of an aluminum halide such as aluminum bromide.

With aluminum bromide as a Lewis acid, we can mix Br₂ with AlBr₃ to give us Br⁺. The presence of Br⁺ is a much better electrophile than Br₂ alone. Bromination is achieved with the help of AlBr₃ (Lewis acid catalyst) as it polarizes the Br-Br bond. The polarization causes the bromine atoms within the Br-Br bond to become more electrophilic. The presence of Br⁺ compared to Br₂ alone is a much better electrophile that can then react with benzene.
As the bromine has now become more electrophilic after activation with a catalyst, an electrophilic attack by the benzene occurs at the terminal bromine of Br-Br-AlBr₃. This allows the other bromine atom to leave with the AlBr₃ as a good leaving group, AlBr₄⁻.

![Chemical reaction diagram]

After the electrophilic attack of bromide to the benzene, the hydrogen on the same carbon as bromine substitutes the carbocation in which resulted from the attack. Hence it being an electrophilic aromatic SUBSTITUTION. Since the by-product aluminum tetrabromide is a strong nucleophile, it pulls of a proton from the Hydrogen on the same carbon as bromine.

![Chemical reaction diagram]

In the end, AlBr₃ was not consumed by the reaction and is regenerated. It serves as our catalyst in the halogenation of benzenes.
Dissociation Energies of Halogens and its Effect on Halogenation of Benzenes

The electrophilic bromination of benzenes is an exothermic reaction. Considering the exothermic rates of aromatic halogenation decreasing down the periodic table in the halogen family, fluorination is the most exothermic and iodination would be the least. Being so exothermic, a reaction of fluorine with benzene is explosive! For iodine, electrophilic iodination is generally endothermic, hence a reaction is often not possible. Similar to bromide, chlorination would require the aid of an activating presence such as aluminum chloride or ferric chloride. The mechanism of this reaction is the same as with bromination of benzene.

Outside links

- [http://www.chemguide.co.uk/mechanism...ogenation.html](http://www.chemguide.co.uk/mechanism...ogenation.html)
- [http://en.wikipedia.org/wiki/Electro...c_halogenation](http://en.wikipedia.org/wiki/Electro...c_halogenation)

References


Problems

1. What reagents would you need to get the given product?

What reagents would you need to get the given product

![Diagram of reaction]

2. What product would result from the given reagents?

![Diagram of reaction]

What product would result from the given reagents?
3. What is the major product given the reagents below?

\[
\text{Br}_2, \text{FeBr}_3 \rightarrow ?
\]

4. Draw the mechanism for the formation of \( \text{Cl}^+ \) from \( \text{AlCl}_3 \) and \( \text{Cl}_2 \)

5. Draw the mechanism of the reaction between \( \text{Cl}^+ \) and benzene.

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**Solutions**

1. \( \text{Cl}_2 \) and \( \text{AlCl}_3 \) or \( \text{Cl}_2 \) and \( \text{FeCl}_3 \)

2. No Reaction

3. 

\[
\begin{align*}
\text{Br} &\rightarrow \text{Cl}^+ \\
\text{AlCl}_3 &\rightarrow \text{Cl}_2
\end{align*}
\]

4. 

\[
\begin{align*}
\text{Cl} &\rightarrow \text{Cl}^+ \\
\text{AlCl}_3 &\rightarrow \text{Cl}_2
\end{align*}
\]
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