Learning Objectives

- To learn about the characteristics of electromagnetic waves. Light, X-Rays, infrared and microwaves are among the types of electromagnetic waves.

Scientists discovered much of what we know about the structure of the atom by observing the interaction of atoms with various forms of radiant, or transmitted, energy, such as the energy associated with the visible light we detect with our eyes, the infrared radiation we feel as heat, the ultraviolet light that causes sunburn, and the x-rays that produce images of our teeth or bones. All these forms of radiant energy should be familiar to you. We begin our discussion of the development of our current atomic model by describing the properties of waves and the various forms of electromagnetic radiation.

Figure \( \PageIndex{1} \): A Wave in Water. When a drop of water falls onto a smooth water surface, it generates a set of waves that travel outward in a circular direction.

Properties of Waves

A wave is a periodic oscillation that transmits energy through space. Anyone who has visited a beach or dropped a stone into a puddle has observed waves traveling through water (Figure \( \PageIndex{1} \)). These waves are produced when wind, a stone, or some other disturbance, such as a passing boat, transfers energy to the water, causing the surface to oscillate up and down as the energy travels outward from its point of origin. As a wave passes a particular point on the surface of the water, anything floating there moves up and down.
Wave properties and electromagnetic radiation.

**Electromagnetic Radiation**

Water waves transmit energy through space by the periodic oscillation of matter (the water). In contrast, energy that is transmitted, or radiated, through space in the form of periodic oscillations of electric and magnetic fields is known as **electromagnetic radiation**. Some forms of electromagnetic radiation are shown in Figure 2. In a vacuum, all forms of electromagnetic radiation—whether microwaves, visible light, or gamma rays—travel at the speed of light \(c\), which turns out to be a fundamental physical constant with a value of \(2.99792458 \times 10^8\) m/s (about \(3.00 \times 10^8\) m/s or \(1.86 \times 10^5\) mi/s). This is about a million times faster than the speed of sound.
Figure $\PageIndex{3}$: The Nature of Electromagnetic Radiation. All forms of electromagnetic radiation consist of perpendicular oscillating electric and magnetic fields.

Because the various kinds of electromagnetic radiation all have the same speed ($c$), they differ in only wavelength and frequency. As shown in Figure $\PageIndex{4}$ and Table $\PageIndex{1}$, the wavelengths of familiar electromagnetic radiation range from $10^1$ m for radio waves to $10^{-12}$ m for gamma rays, which are emitted by nuclear reactions. By replacing $v$ with $c$ in Equation \ref{6.1.1a}, we can show that the frequency of electromagnetic radiation is inversely proportional to its wavelength:

$$c = \lambda \nu \quad \text{or} \quad \nu = \frac{c}{\lambda} \label{6.1.2}$$

For example, the frequency of radio waves is about $10^8$ Hz, whereas the frequency of gamma rays is about $10^{20}$ Hz. Visible light, which is electromagnetic radiation that can be detected by the human eye, has wavelengths between about $7 \times 10^{-7}$ m (700 nm, or $4.3 \times 10^{14}$ Hz) and $4 \times 10^{-7}$ m (400 nm, or $7.5 \times 10^{14}$ Hz). Note that when frequency increases, wavelength decreases; $c$ being a constant stays the same. Similarly, when frequency decreases, the wavelength increases.

Figure $\PageIndex{4}$: The Electromagnetic Spectrum. (a) This diagram shows the wavelength and frequency ranges of electromagnetic radiation. The visible portion of the electromagnetic spectrum is the narrow region with wavelengths between about 400 and 700 nm. (b) When white light is passed through a prism, it is split into light of different wavelengths, whose colors correspond to the visible spectrum.

Within the visible range our eyes perceive radiation of different wavelengths (or frequencies) as light of different colors, ranging from red to violet in order of decreasing wavelength. The components of white light—a mixture of all the frequencies of visible light—can be separated by a prism, as shown in part (b) in Figure $\PageIndex{4}$. A similar phenomenon creates a rainbow, where water droplets suspended in the air act as tiny prisms.

Table $\PageIndex{1}$: Common Wavelength Units for Electromagnetic Radiation

<table>
<thead>
<tr>
<th>Unit</th>
<th>Symbol</th>
<th>Wavelength (m)</th>
<th>Type of Radiation</th>
</tr>
</thead>
<tbody>
<tr>
<td>picometer</td>
<td>pm</td>
<td>$10^{-12}$</td>
<td>gamma ray</td>
</tr>
</tbody>
</table>
As you will soon see, the energy of electromagnetic radiation is directly proportional to its frequency and inversely proportional to its wavelength:

\[
E \propto u \quad \text{(Equation 6.1.3)}
\]
\[
E \propto \frac{1}{\lambda} \quad \text{(Equation 6.1.4)}
\]

Whereas visible light is essentially harmless to our skin, ultraviolet light, with wavelengths of \( \leq 400 \text{ nm} \), has enough energy to cause severe damage to our skin in the form of sunburn. Because the ozone layer of the atmosphere absorbs sunlight with wavelengths less than 350 nm, it protects us from the damaging effects of highly energetic ultraviolet radiation.

The energy of electromagnetic radiation increases with increasing frequency and decreasing wavelength.

Example (PageIndex{1})): Wavelength of Radiowaves

Your favorite FM radio station, WXYZ, broadcasts at a frequency of 101.1 MHz. What is the wavelength of this radiation?

**Given:** frequency

**Asked for:** wavelength

**Strategy:**

Substitute the value for the speed of light in meters per second into Equation (ref{6.1.2}) to calculate the wavelength in meters.

**Solution:**

From Equation (ref{6.1.2}), we know that the product of the wavelength and the frequency is the speed of the wave, which for electromagnetic radiation is \( 2.998 \times 10^8 \text{ m/s} \):
Thus the wavelength $\lambda$ is given by

\[
\lambda = \frac{c}{u} = \left( \frac{2.988 \times 10^8 \text{ m/s}}{101.1 \text{ MHz}} \right) \left( \frac{1 \text{ MHz}}{10^6 \text{ s}^{-1}} \right) = 2.965 \text{ m}
\]

Exercise \(\PageIndex{1}\)

As the police officer was writing up your speeding ticket, she mentioned that she was using a state-of-the-art radar gun operating at 35.5 GHz. What is the wavelength of the radiation emitted by the radar gun?

Answer

8.45 mm

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**Summary**

Understanding the electronic structure of atoms requires an understanding of the properties of waves and electromagnetic radiation. A wave is a periodic oscillation by which energy is transmitted through space. All waves are periodic, repeating regularly in both space and time. Waves are characterized by several interrelated properties: wavelength ($\lambda$), the distance between successive waves; frequency ($f$), the number of waves that pass a fixed point per unit time; speed ($v$), the rate at which the wave propagates through space; and amplitude, the magnitude of the oscillation about the mean position. The speed of a wave is equal to the product of its wavelength and frequency.

Electromagnetic radiation consists of two perpendicular waves, one electric and one magnetic, propagating at the speed of light ($c$). Electromagnetic radiation is radiant energy that includes radio waves, microwaves, visible light, x-rays, and gamma rays, which differ in their frequencies and wavelengths.

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**Contributors and Attributions**

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