One of the largest and most diverse classes of reactions is composed of nucleophilic additions to a carbonyl group. Conjugation of a double bond to a carbonyl group transmits the electrophilic character of the carbonyl carbon to the beta-carbon of the double bond. These conjugated carbonyl are called enones or $\alpha, \beta$ unsaturated carbonyls. A resonance description of this transmission is shown below.

From this formula it should be clear that nucleophiles may attack either at the carbonyl carbon, as for any aldehyde, ketone or carboxylic acid derivative, or at the beta-carbon. These two modes of reaction are referred to as 1,2-addition and 1,4-addition respectively. A 1,4-addition is also called a conjugate addition.

**Basic reaction of 1,2 addition**

Here the nucleophile adds to the carbon which is in the one position. The hydrogen adds to the oxygen which is in the two position.

![1,2 addition reaction](image)

**Basic reaction of 1,4 addition**

In 1,4 addition the Nucleophile is added to the carbon $\beta$ to the carbonyl while the hydrogen is added to the carbon $\alpha$ to the carbonyl.

**Mechanism for 1,4 addition**

1) Nucleophilic attack on the carbon $\beta$ to the carbonyl
2) Proton Transfer

Here we can see why this addition is called 1,4. The nucleophile bonds to the carbon in the one position and the hydrogen adds to the oxygen in the four position.

3) Tautomerization

Going from reactant to products simplified
1,2 vs. 1,4 addition

Whether 1,2 or 1,4-addition occurs depends on multiple variables but mostly it is determined by the nature of the nucleophile. During the addition of a nucleophile there is a competition between 1,2 and 1,4 addition products. If the nucleophile is a strong base, such as Grignard reagents, both the 1,2 and 1,4 reactions are irreversible and therefore are under kinetic control. Since 1,2-additions to the carbonyl group are fast, we would expect to find a predominance of 1,2-products from these reactions.

If the nucleophile is a weak base, such as alcohols or amines, then the 1,2 addition is usually reversible. This means the competition between 1,2 and 1,4 addition is under thermodynamic control. In this case 1,4-addition dominates because the stable carbonyl group is retained.

Nucleophiles which add 1,4 to α, β unsaturated carbonyls

Water

![Diagram of the reaction between water and an α, β unsaturated carbonyl group showing 1,4-addition to form a hydroxy group.]

Alcohols

![Diagram of the reaction between an alcohol and an α, β unsaturated carbonyl group showing 1,4-addition to form an ether group.]

Thiols

![Diagram of the reaction between a thiol and an α, β unsaturated carbonyl group showing 1,4-addition to form a sulfide group.]

1° Amines

![Diagram of the reaction between a primary amine and an α, β unsaturated carbonyl group showing 1,4-addition to form an amide group.]

Another important reaction exhibited by organometallic reagents is metal exchange. Organolithium reagents react with cuprous iodide to give a lithium dimethylcopper reagent, which is referred to as a Gilman reagent. Gilman reagents are a source of carbanion like nucleophiles similar to Grignard and Organo lithium reagents. However, the reactivity of organocuprate reagents is slightly different and this difference will be exploited in different situations. In the case of $\alpha$, $\beta$ unsaturated carbonyls organocuprate reagents allow for an 1,4 addition of an alkyl group. As we will see later Grignard and Organolithium reagents add alkyl groups 1,2 to $\alpha$, $\beta$ unsaturated carbonyls.
Organocuprate reagents are made from the reaction of organolithium reagents and $\text{CuI}$

$$[2 \text{RLi} + \text{CuI} \rightarrow \text{R}_2\text{CuLi} + \text{LiI}]$$

This acts as a source of $R^-$

$$[2 \text{CH}_3\text{Li} + \text{CuI} \rightarrow (\text{CH}_3)_2\text{CuLi} + \text{LiI}]$$

**Example**

Nucleophiles which add 1,2 to $\alpha, \beta$ unsaturated carbonyls

**Metal Hydrides**

**Grignard Reagents**
Organolithium Reagents

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