The following set of experiments introduce students to the applications of titrimetry. Experiments are grouped into four categories based on the type of reaction (acid–base, complexation, redox, and precipitation). Additional experiments emphasizing potentiometric electrodes are found in Chapter 11.

**Acid–Base Titrimetry**


**Complexation Titrimetry**

- Fulton, R.; Ross, M.; Schroeder, K. “Spectrophotometric Titration of a Mixture of Calcium and Magnesium,” *J.


### RedoxTitration


### PrecipitationTitration


For a general history of titrimetry, see the following sources.


The use of weight instead of volume as a signal for titrimetry is reviewed in the following paper.


A more thorough discussion of non-aqueous titrations, with numerous practical examples, is provided in the following text.

- Fritz, J. S. *Acid-Base Titrations in Nonaqueous Solvents*; Allyn and Bacon, Boston; 1973.

The sources listed below provides more details on the how potentiometric titration data may be used to calculate
equilibrium constants.


The following provides additional information about Gran plots.


The following provide additional information about calculating or sketching titration curves.


For a complete discussion of the application of complexation titrimetry see the texts and articles listed below.


A good source for additional examples of the application of all forms of titrimetry is