Skills to Develop

- To calculate a mass-energy balance and a nuclear binding energy.
- To understand the differences between nuclear fission and fusion.

Nuclear reactions, like chemical reactions, are accompanied by changes in energy. The energy changes in nuclear reactions, however, are enormous compared with those of even the most energetic chemical reactions. In fact, the energy changes in a typical nuclear reaction are so large that they result in a measurable change of mass. In this section, we describe the relationship between mass and energy in nuclear reactions and show how the seemingly small changes in mass that accompany nuclear reactions result in the release of enormous amounts of energy.

Mass–Energy Balance

The relationship between mass \( m \) and energy \( E \) is expressed in the following equation:

\[
E = mc^2 \label{Eq1}
\]

where

- \( c \) is the speed of light \((2.998 \times 10^8 \text{ m/s})\), and
- \( E \) and \( m \) are expressed in units of joules and kilograms, respectively.

Albert Einstein first derived this relationship in 1905 as part of his special theory of relativity: the mass of a particle is directly proportional to its energy. Thus according to Equation \(\ref{Eq1}\), every mass has an associated energy, and similarly, any reaction that involves a change in energy must be accompanied by a change in mass. This implies that all exothermic reactions should be accompanied by a decrease in mass, and all endothermic reactions should be accompanied by an increase in mass. Given the law of conservation of mass, how can this be true? The solution to this apparent contradiction is that chemical reactions are indeed accompanied by changes in mass, but these changes are simply too small to be detected. As you may recall, all particles exhibit wavelike behavior, but the wavelength is inversely proportional to the mass of the particle (actually, to its momentum, the product of its mass and velocity). Consequently, wavelike behavior is detectable only for particles with very small masses, such as electrons. For example, the chemical equation for the combustion of graphite to produce carbon dioxide is as follows:

\[
\text{C(graphite)} + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) \hspace{5mm} \Delta H^\circ= -393.5 \text{ kJ/mol} \label{Eq2}
\]

Combustion reactions are typically carried out at constant pressure, and under these conditions, the heat released or absorbed is equal to \( \Delta H \). When a reaction is carried out at constant volume, the heat released or absorbed is equal to \( \Delta E \). For most chemical reactions, however, \( \Delta E = \Delta H \). If we rewrite Einstein’s equation as

\[
\Delta E = (\Delta m)c^2 \label{Eq3}
\]

we can rearrange the equation to obtain the following relationship between the change in mass and the change in energy:
\[ \Delta m = \frac{\Delta E}{c^2} \]

Because 1 J = 1 (kg•m²)/s², the change in mass is as follows:
\[ \Delta m = \frac{-393.5\, \text{kJ/mol}}{(2.998\times10^8\, \text{m/s})^2} = -4.38\times10^{-12} \text{ kg/mol} \]

This is a mass change of about 3.6 \times 10^{-10} g/g carbon that is burned, or about 100-millionths of the mass of an electron per atom of carbon. In practice, this mass change is much too small to be measured experimentally and is negligible.

In contrast, for a typical nuclear reaction, such as the radioactive decay of \(^{14}\text{C}\) to \(^{14}\text{N}\) and an electron (a \(\beta\) particle), there is a much larger change in mass:
\[ ^{14}\text{C} \rightarrow ^{14}\text{N} + ^{0}_{-1}\beta \]

We can use the experimentally measured masses of subatomic particles and common isotopes given in Table 20.1 to calculate the change in mass directly. The reaction involves the conversion of a neutral \(^{14}\text{C}\) atom to a positively charged \(^{14}\text{N}\) ion (with six, not seven, electrons) and a negatively charged \(\beta\) particle (an electron), so the mass of the products is identical to the mass of a neutral \(^{14}\text{N}\) atom. The total change in mass during the reaction is therefore the difference between the mass of a neutral \(^{14}\text{N}\) atom (14.003074 amu) and the mass of a \(^{14}\text{C}\) atom (14.003242 amu):
\[ \begin{align} \Delta m &= \text{mass}_{\text{products}} - \text{mass}_{\text{reactants}} \\ &= 14.003074\, \text{amu} - 14.003242\, \text{amu} = -0.000168\, \text{amu} \end{align} \]

The energy released in this nuclear reaction is more than 100,000 times greater than that of a typical chemical reaction, even though the decay of \(^{14}\text{C}\) is a relatively low-energy nuclear reaction.

\[ \Delta E = (\Delta m)c^2 = (-1.68\times10^{-7}\, \text{kg})(2.998\times10^8\, \text{m/s})^2 = -1.51\times10^{10}\, \text{J} = -1.51\times10^7\, \text{kJ} \]

Nuclear Binding Energies

We have seen that energy changes in both chemical and nuclear reactions are accompanied by changes in mass. Einstein’s equation, which allows us to interconvert mass and energy, has another interesting consequence: The mass of an atom is always less than the sum of the masses of its component particles. The only exception to this rule is
hydrogen-1 ($^{1}$H), whose measured mass of 1.007825 amu is identical to the sum of the masses of a proton and an electron. In contrast, the experimentally measured mass of an atom of deuterium ($^{2}$H) is 2.014102 amu, although its calculated mass is 2.016490 amu:

$$m_{^{2}\text{H}} = m_{\text{neutron}} + m_{\text{proton}} + m_{\text{electron}}$$
$$= 1.008665\text{ amu} + 1.007276\text{ amu} + 0.000549\text{ amu} = 2.016490\text{ amu}$$

The difference between the sum of the masses of the components and the measured atomic mass is called the mass defect of the nucleus. Just as a molecule is more stable than its isolated atoms, a nucleus is more stable (lower in energy) than its isolated components. Consequently, when isolated nucleons assemble into a stable nucleus, energy is released. According to Equation \ref{Eq4}, this release of energy must be accompanied by a decrease in the mass of the nucleus.

The amount of energy released when a nucleus forms from its component nucleons is the nuclear binding energy (Figure \ref{Fig1}). In the case of deuterium, the mass defect is 0.002388 amu, which corresponds to a nuclear binding energy of 2.22 MeV for the deuterium nucleus. Because the magnitude of the mass defect is proportional to the nuclear binding energy, both values indicate the stability of the nucleus.

![Figure 1: Nuclear Binding Energy in Deuterium. The mass of a $^{2}$H atom is less than the sum of the masses of a proton, a neutron, and an electron by 0.002388 amu; the difference in mass corresponds to the nuclear binding energy. The larger the value of the mass defect, the greater the nuclear binding energy and the more stable the nucleus.](image)

Mass defect = 0.002388 amu

Not all nuclei are equally stable. Chemists describe the relative stability of different nuclei by comparing the binding energy per nucleon, which is obtained by dividing the nuclear binding energy by the mass number (A) of the nucleus. As shown in Figure \ref{Fig2}, the binding energy per nucleon increases rapidly with increasing atomic number until about Z
= 26, where it levels off to about 8–9 MeV per nucleon and then decreases slowly. The initial increase in binding energy is not a smooth curve but exhibits sharp peaks corresponding to the light nuclei that have equal numbers of protons and neutrons (e.g., $^4\text{He}$, $^{12}\text{C}$, and $^{16}\text{O}$). As mentioned earlier, these are particularly stable combinations.

**Figure \(\PageIndex{2}\):** The Curve of Nuclear Binding Energy. This plot of the average binding energy per nucleon as a function of atomic number shows that the binding energy per nucleon increases with increasing atomic number until about $Z = 26$, levels off, and then decreases. The sharp peaks correspond to light nuclei that have equal numbers of protons and neutrons.

Because the maximum binding energy per nucleon is reached at $^{56}\text{Fe}$, all other nuclei are thermodynamically unstable with regard to the formation of $^{56}\text{Fe}$. Consequently, heavier nuclei (toward the right in Figure \(\PageIndex{2}\)) should spontaneously undergo reactions such as alpha decay, which result in a decrease in atomic number. Conversely, lighter elements (on the left in Figure \(\PageIndex{2}\)) should spontaneously undergo reactions that result in an increase in atomic number. This is indeed the observed pattern.

Heavier nuclei spontaneously undergo nuclear reactions that decrease their atomic number. Lighter nuclei spontaneously undergo nuclear reactions that increase their atomic number.

### Summary

Unlike a chemical reaction, a nuclear reaction results in a significant change in mass and an associated change of energy, as described by Einstein’s equation. Nuclear reactions are accompanied by large changes in energy, which result in detectable changes in mass. The change in mass is related to the change in energy according to Einstein’s equation: $\Delta E = (\Delta m)c^2$. Large changes in energy are usually reported in kiloelectronvolts or megaelectronvolts (thousands or millions of electronvolts). With the exception of $^1\text{H}$, the experimentally determined mass of an atom is always less than the sum of the masses of the component particles (protons, neutrons, and electrons) by an amount called the mass defect of the nucleus. The energy corresponding to the mass defect is the nuclear binding energy, the amount of energy released when a nucleus forms from its component particles. In nuclear fission, nuclei split into lighter nuclei with an accompanying release of multiple neutrons and large amounts of energy. The critical mass is the minimum mass required to support a self-sustaining nuclear chain reaction. Nuclear fusion is a process in which two light nuclei combine to produce a heavier nucleus plus a great deal of energy.