The strong oxidizing agents most commonly encountered when discussing alcohols are KMnO4 and Chromic acid

Potassium Permanganate

Potassium permanganate, KMnO₄, is a powerful oxidizing agent, and has many uses in organic chemistry.

Introduction

Of all the oxidizing agents discussed in organic chemistry textbooks, potassium permanganate, KMnO₄, is probably the most common, and also the most applicable. As will be shown below, KMnO₄ can be utilized to oxidize a wide range of organic molecules. The products that are obtained can vary depending on the conditions, but because KMnO₄ is such a strong oxidizing agent, the final products are often carboxylic acids.

The half-reaction and oxidation potential

Mn(VII) is reduced under acidic conditions to Mn(IV) or Mn(II) according to the half-reactions shown below, with the indicated cell potentials

\[
\text{MnO}_4^- + 4H^+ + 3e^- \rightarrow \text{MnO}_2 + 2H_2O; \ E^o = 1.68 \ V
\]

\[
\text{MnO}_4^- + 8H^+ + 5e^- \rightarrow \text{Mn}^{2+} + 4H_2O; \ E^o = 1.5 \ V
\]

\[
\text{MnO}_4^- + 2H_2O + 3e^- \rightarrow \text{Mn}^{2+} + 4OH^-; \ E^o = 0.6 \ V
\]

General Reactivity with Organic Molecules

KMnO₄ is able to oxidize carbon atoms if they contain sufficiently weak bonds, including

1. Carbon atoms with \(\text{\pi}\) bonds, as in alkenes and alkynes
2. Carbon atoms with weak C-H bonds, such as
   - C-H bonds in the alpha-positions of substituted aromatic rings
   - C-H bonds in carbon atoms containing C-O bonds, including alcohols and aldehydes
3. Carbons with exceptionally weak C-C bonds such as
   - C-C bonds in a glycol
   - C-C bonds next to an aromatic ring AND an oxygen

KMnO₄ also oxidizes phenol to para-benzoquinone.

Example
Examples of carbons that are not oxidized

a. Aliphatic carbons (except those alpha to an aromatic ring, as above)
b. Aromatic carbons (except phenol, as above)
c. Carbons without a C-H bond, except as in (3) above

Exhaustive oxidation of organic molecules by KMnO$_4$ will proceed until the formation of carboxylic acids. Therefore, alcohols will be oxidized to carbonyls (aldehydes and ketones), and aldehydes (and some ketones, as in (3) above) will be oxidized to carboxylic acids.

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**Reactions with Specific Functional Groups**

Using the principles above, we expect KMnO$_4$ to react with alkenes, alkynes, alcohols, aldehydes and aromatic side chains. Examples are provided below. It is easiest to start at the top.

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**Aldehydes**

Aldehydes RCHO are readily oxidized to carboxylic acids.

\[
\text{CH}_2\text{CN/H}_2\text{O} + \text{K MnO}_4 \rightarrow \text{CH}_2\text{COOH}
\]

Unless great efforts are taken to maintain a neutral pH, KMnO$_4$ oxidations tend to occur under basic conditions. In fact, the most effective conditions for aldehyde oxidation by KMnO$_4$ involves t-butanol as solvent with a NaH$_2$PO$_4$ buffer.$^2$ The reactions above are deliberately not balanced equations. Balancing the reactions would involve using the methods learned in general chemistry, requiring half reactions for all processes.

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**Alcohols**

Primary alcohols such as octan-1-ol can be oxidized efficiently by KMnO$_4$, in the presence of basic copper salts.$^3$ However, the product is predominantly octanoic acid, with only a small amount of aldehyde, resulting from overoxidation.

\[
\text{CH}_3\text{(CH}_3\text{)}\text{OH} \xrightarrow{\text{K MnO}_4/\text{CuSO}_4\text{H}_2\text{O}/\text{KOH}} \text{CH}_3\text{(CH}_3\text{)}\text{CHO} \rightarrow \text{CH}_3\text{(CH}_3\text{)}\text{CO}_2\text{H}
\]

Although overoxidation is less of a problem with secondary alcohols, KMnO$_4$ is still not considered generally well-suited for conversions of alcohols to aldehydes or ketones.

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**Alkenes**$^4$

Under mild conditions, potassium permanganate can effect conversion of alkenes to glycols. It is, however, capable of further oxidizing the glycol with cleavage of the carbon-carbon bond, so careful control of the reaction conditions is
necessary. A cyclic manganese diester is an intermediate in these oxidations, which results in glycols formed by syn addition.

\[
\text{Cyclic manganese diester} \xrightarrow{\text{KMnO}_4} \text{Glycol}
\]

With addition of heat and/or more concentrated KMnO\(_4\), the glycol can be further oxidized, cleaving the C-C bond.

\[
\text{Glycol} \xrightarrow{\text{KMnO}_4, \text{acetone}} \text{Carboxylic acid}
\]

More substituted olefins will terminate at the ketone

\[
\text{Olefins} \xrightarrow{\text{KMnO}_4, \text{acetone}} \text{Ketone}
\]

Oxidative cleavage of the diol can be carried out more mildly by using IO\(_4\) as the oxidant.

The cleavage of alkenes to ketones/carboxylic acids can be used to determine the position of double bonds in organic molecules. \(^5\)

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### Alkynes\(^4\)

Instead of bis-hydroxylation that occurs with alkenes, permanganate oxidation of alkynes initially leads to the formation of diones.

\[
\text{PhC}≡\text{CH}_2\text{CH}_2\text{CH}_3 \xrightarrow{\text{KMnO}_4, \text{R}_4\text{N}^+, \text{CH}_2\text{Cl}_2} \text{PhC}≡\text{C}\text{H}_2\text{CH}_2\text{CH}_3
\]

Under harsher conditions, the dione is cleaved to form two carboxylic acids.

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### Aromatic side-chains\(^6\)

Treatment of an alkylbenzene with potassium permanganate results in oxidation to give the benzoic acid.

\[
\text{Conversion of alkylbenzenes to benzoic acids using KMnO}_4
\]

Alkybenzene

Benzoic acid

Key bonds formed: C-O\(_i\), C-O, C-OH

Key bonds broken: C-H

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Note
The position directly adjacent to an aromatic group is called the “benzylic” position.

The reaction only works if there is at least one hydrogen attached to the carbon. However, if there is at least one hydrogen, the oxidation proceeds all the way to the carboxylic acid.

**Examples:**

![Chemical structures](image)

Notes: Note that in example 2 the extra carbons are cleaved to give the same product as in example 1. And in example 3, two benzoic acids are formed. Finally, when no hydrogens are present on the benzylic carbon, no reaction occurs (example 4).

The oxidation of alkyl side-chains to form benzoic acids was historically used in qualitative analysis to determine the positions of alkyl groups in substituted aromatic systems. Alkyl-substituted rings can be converted to poly-acids, which can be distinguished on the basis of their pKas.

**Additional Reading**

- Oxidation by Chromic Acid (H$_2$CrO$_4$)
- Ozonolysis
- Oxidative cleavage of double bonds
- Oxidation of alkenes
Chromic Acid

One of the reagents that is commonly used for oxidation in organic chemistry is chromic acid. This reagent is straightforward to use once deciphered. However, there are a vast number of different ways that textbooks (and instructors) show it being used in reactions.

Chromic acid, $\text{H}_2\text{CrO}_4$, is a strong acid and is a reagent for oxidizing alcohols to ketones and carboxylic acids. For reasons primarily concerning safety and convenience, chromic acid tends to be produced in a reaction vessel as needed (through the addition of acid to a source of chromium), rather than being dispensed from a bottle.

This is where the trouble begins. Choosing a source of chromium to produce $\text{H}_2\text{CrO}_4$ is a lot like choosing a favorite brand of bottled water. Beyond the packaging, they’re pretty much all the same. Depending on which textbook or instructor you have, however, you might see several different ways to do this, and it can be very confusing.

The key point is that $\text{Na}_2\text{CrO}_4$ (sodium chromate), $\text{Na}_2\text{Cr}_2\text{O}_7$ (sodium dichromate), $\text{KCrO}_4$ (potassium chromate), $\text{K}_2\text{Cr}_2\text{O}_7$ (potassium dichromate), and $\text{CrO}_3$ (chromium trioxide) are alike in one crucial manner: when they are combined with aqueous acid, each of them forms $\text{H}_2\text{CrO}_4$, and ultimately it is $\text{H}_2\text{CrO}_4$ that is involved in the important chemistry. Unfortunately I rarely see this point explained in textbooks. I remember this causing some confusion for me when I took the course. The K or Na ions present are just spectators.

Once $\text{H}_2\text{CrO}_4$ is formed, its reactions are pretty straightforward: it converts primary alcohols (and aldehydes) to
carboxylic acids and secondary alcohols to ketones. Once deciphered, chromic acid is a fairly straightforward reagent. It oxidizes primary alcohols to carboxylic acids and secondary alcohols to ketones. It will also oxidize aldehydes to ketones.

It does this through the addition of the alcohol oxygen to chromium, which makes it a good leaving group; a base (water being the most likely culprit) can then remove a proton from the carbon, forming a new $\pi$ bond and breaking the O-Cr bond.

Due to its high toxicity, chromic acid tends to find very little use in the organic chemistry laboratory outside of undergrad labs, and there are far more useful reagents available for performing these transformations.
Contributors

- James Ashenhurst ([MasterOrganicChemistry.com](http://MasterOrganicChemistry.com))

Additional Resources

*Web Pages*

[Jones reagent](http://Jones reagent)