The collisional cross section is an "effective area" that quantifies the likelihood of a scattering event when an incident species strikes a target species. In a hard object approximation, the cross section is the area of the conventional geometric cross section. The collisional cross sections typically denoted \( \sigma \) and measured in units of area.

**Introduction**

Atoms and molecules can move around in space and bump into each other. If certain conditions of the collision are met, a chemical reaction occurs and a product forms. Sometimes, however, particles may get extremely close to each other but do not strike. We can use the collisional cross section to find how large the distance between two particles must be in order for a collision to occur.

A few assumptions must be made:

- All particles travel through space linearly
- All particles are hard spheres
- Only two particles are involved in the collision

The collisional cross section is defined as the area around a particle in which the center of another particle must be in order for a collision to occur. In the image below, the area within the black circle is molecule A's collisional cross section. That area can change depending on the size of the two particles involved.

![Collisional Cross Section Diagram](image)

Figure (\( \PageIndex{1} \)): Distance is greater than \( \langle A \rangle \) and \( \langle B \rangle \) radii, so no collision occurs.

The collisional cross section \( \sigma_{AA} \) between molecule \( \langle A \rangle \) and molecule \( \langle A \rangle \) can be calculated using the following equation:

\[
\sigma_{AB} = \pi(r_A + r_B)^2
\]

Collision occurs when the distance between the center of the two reactant molecules is less than the sum of the radii of these molecules, as shown in Figure (\( \PageIndex{2} \)). The collisional cross section describes the area around a single reactant. For a collisional reaction to occur, the center of one reactant must be within the collisional cross section of a corresponding reactant.
Figure \(\PageIndex{2}\): Distance is less than \((A)\) and \((B)\) radii, so no collision occurs.

It is first assumed that all particles, whether it be an atom or a molecule, are hard spheres. Particle A must come in contact with Particle B in order for a collision to occur. Particle B can approach particle A from any direction; thus, consider a circle with radius \(r\):

\[
\text{Area of a Circle} = \pi r^2
\]

Assume that the two particles involved in the collision are the same in size and have the same radius. The furthest distance the two centers can be and still have a collision is \((2r)\). Substitute \((r)\) with \((2r)\):

\[
\sigma_{AA} = \pi (2r)^2 \quad \text{(Same Eq)}
\]

We will define this area as the collisional cross section. Anytime the center of another particle is within this area, there will be parts of the two particles that will overlap, touch, and cause a collision.

Example \(\PageIndex{1}\): Hydrogen/Hydrogen Collisions

Consider the following reaction: \(H + H \rightarrow H_2\) The radius of hydrogen is \(5.3 \times 10^{-11}\; m\). What is the collisional cross section for this reaction?

Solution

Use Equation \(\ref{SameEq}\) with the given atomic radius for hydrogen atoms:

\[
\text{Collisional Cross Section} = \pi (2r)^2 = \pi [(2)(5.3 \times 10^{-11})^2] = 3.53 \times 10^{-20} m^2
\]

Although the collisional cross section of a particle can be calculated, it is usually not used on its own (Table \(\PageIndex{1}\)). Instead, it is a component of more complex theories such as collision frequency and collision theory.

Figure \(\PageIndex{1}\): Table of Cross-Sections for Common Gases

<table>
<thead>
<tr>
<th>Molecule</th>
<th>Cross-Section (nm(^2))</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ar</td>
<td>0.36</td>
</tr>
<tr>
<td>C(_2)H(_4)</td>
<td>0.64</td>
</tr>
</tbody>
</table>
### Table: Molecules and Cross-Sections (nm²)

<table>
<thead>
<tr>
<th>Molecule</th>
<th>Cross-Section (nm²)</th>
</tr>
</thead>
<tbody>
<tr>
<td>C₆H₆</td>
<td>0.88</td>
</tr>
<tr>
<td>CH₄</td>
<td>0.46</td>
</tr>
<tr>
<td>Cl₂</td>
<td>0.93</td>
</tr>
<tr>
<td>CO₂</td>
<td>0.52</td>
</tr>
<tr>
<td>H₂</td>
<td>0.27</td>
</tr>
<tr>
<td>He</td>
<td>0.21</td>
</tr>
<tr>
<td>N₂</td>
<td>0.43</td>
</tr>
<tr>
<td>Ne</td>
<td>0.24</td>
</tr>
<tr>
<td>O₂</td>
<td>0.40</td>
</tr>
<tr>
<td>SO₂</td>
<td>0.58</td>
</tr>
</tbody>
</table>

Now what if the particles were of different size and different radii? The \(2r\) term in Equation \(\ref{SameEq}\) is really the sum of the radius of each molecule (i.e., \(2r = r + r\)). However, if the colliding molecules have differing sizes (e.g., \(r_A\) and \(r_B\)) for the radius of particle A and B, respectively, then \(2r\) in Equation \(\ref{SameEq}\) is substituted with \(r_A + r_B\):

\[
\sigma_{AB} = \pi (r_A + r_B)^2
\]

**Example \(\PageIndex{1}\): Hydrogen/Fluorine Collisions**

What is the collisional cross section for this reaction?

\[
H + F \rightarrow HF
\]

The radius of fluorine atom is \(4.2 \times 10^{-11}\) m.

**Solution**

\[
\text{Collisional Cross-Section} = \pi (r_A + r_B)^2 = \pi ([5.3 \times 10^{-11}] + [4.3 \times 10^{-11}])^2 = 2.90 \times 10^{-20} \text{ m}^2
\]

**Problems**

1. \(H_2 + O_2 \rightarrow H_2O\). The radius of oxygen is \(4.8 \times 10^{-11}\) m. What is the collisional cross section for this reaction? (Hint: Assume that the radius of a molecule is just the sum of its atoms.)
2. \( \text{N}_2 + \text{O}_2 \rightarrow \text{N}_2\text{O} \). The radius of nitrogen is \( 5.6 \times 10^{-11} \) m. If the distance between the two centers is \( 2.00 \times 10^{-19} \) m, is there a collision between the two molecules?

3. \( \text{F} + \text{F} \rightarrow \text{F}_2 \) The center of the two atoms are \( 3.50 \times 10^{-20} \) m apart from each other. How much closer do the centers have to be in order for a collision to occur?

**Answers**

1. The radius of \( \text{H}_2 \) is \( 2(\text{r}_\text{H})=1.06 \times 10^{-10} \) m. The radius of \( \text{O}_2 \) is \( 2(\text{r}_\text{O})=9.6 \times 10^{-11} \) m. 

   \[ \text{Collisional Cross Section} = \pi(\text{r}_A+\text{r}_B)^2 = \pi((1.06 \times 10^{-10})+(9.6 \times 10^{-11}))^2 = 1.28 \times 10^{-19} \]

2. Yes, there is a collision. The radius of \( \text{N}_2 \) is \( 2(\text{r}_\text{N})=1.12 \times 10^{-10} \) m. The radius of \( \text{O}_2 \) is \( 2(\text{r}_\text{O})=9.6 \times 10^{-11} \) m.

   \[ \text{Collisional Cross Section} = \pi(\text{r}_A+\text{r}_B)^2 = \pi((1.12 \times 10^{-10})+(9.6 \times 10^{-11}))^2 = 1.36 \times 10^{-19} \]

   The furthest the two molecules can be and still get a collision. Since \( 2.00 \times 10^{-19} \) m is larger than that distance, the center of one molecule is not in the collisional cross section of the other molecule. Therefore, no collision occurs. The molecules are too far apart.

3. \( \text{Collisional Cross Section} = \pi(2\text{r})^2 = \pi((2)(4.2 \times 10^{-11}))^2 = 2.22 \times 10^{-20} \]

   Because the molecules are \( 3.50 \times 10^{-20} \) m apart, the center of one F is not in the Collisional Cross Section of the other F. They must be closer. \( (3.50 \times 10^{-20})-2.22 \times 10^{-20}=1.28 \times 10^{-20} \]

   The molecules must be at least \( 1.28 \times 10^{-20} \) m closer.

**References**


**Contributors and Attributions**

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