Learning Objective

- predict the products/specify the reagents for dihydroxylation of alkenes

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**Epoxide ring-opening reactions - $S_{N}1$ vs. $S_{N}2$, regioselectivity, and stereoselectivity**

The nonenzymatic ring-opening reactions of epoxides provides an opportunity to review the nucelophilic substitution mechanisms. Ring-opening reactions can proceed by either $S_{N}2$ or $S_{N}1$ mechanisms, depending on the nature of the epoxide and on the reaction conditions. If the epoxide is asymmetric, the structure of the product will vary according to which mechanism dominates. When an asymmetric epoxide undergoes solvolysis in basic methanol, ring-opening occurs by an $S_{N}2$ mechanism, and the less substituted carbon reacts with the nucleophile under steric considerations and produces product B in the example below.

![Basic ring-opening](image1)

Conversely, when solvolysis occurs in acidic methanol, the reaction occurs by a mechanism with substantial $S_{N}1$ character, and the more substituted carbon reacts with the nucleophile under electrostatic considerations and produces product A in the example below.

![Acidic ring-opening](image2)

These are both good examples of **regioselective reactions**. In a regioselective reaction, two (or more) different constitutional isomers are possible as products, but one is formed preferentially (or
sometimes exclusively).

Let us examine the basic, \(S_N2\) case first. The leaving group is an alkoxide anion, because there is no acid available to protonate the oxygen prior to ring opening. An alkoxide is a poor leaving group, and thus the ring is unlikely to open without a 'push' from the nucleophile.

![Diagram of \(S_N2\) mechanism]

The nucleophile itself is potent: a deprotonated, negatively charged methoxide ion. When a nucleophilic substitution reaction involves a poor leaving group and a powerful nucleophile, it is very likely to proceed by an \(S_N2\) mechanism.

What about the electrophile? There are two electrophilic carbons in the epoxide, but the best target for the nucleophile in an \(S_N2\) reaction is the carbon that is \textit{least hindered}. This accounts for the observed regiochemical outcome. Like in other \(S_N2\) reactions, nucleophilic reactions take place with backside orientation relative to the leaving group, resulting in inversion at the electrophilic carbon.

Probably the best way to depict the acid-catalyzed epoxide ring-opening reaction is as a hybrid, or cross, between an \(S_N2\) and \(S_N1\) mechanism. First, the oxygen is protonated, creating a good leaving group (step 1 below). Electrostatic considerations have greater importance with a protonated intermediate. As the carbon-oxygen bond begins to break (step 2), positive charge builds on the more substituted carbon with greater carbocation stability.
Unlike in an S_N1 reaction, the nucleophile reacts with the electrophilic carbon (step 3) before a complete carbocation intermediate has a chance to form.

Reaction takes place preferentially from the backside (like in an S_N2 reaction) because the carbon-oxygen bond is still to some degree in place, and the oxygen blocks reaction from the front side. Notice, however, the regiochemical outcome is different from the base-catalyzed reaction. In the acid-catalyzed process, the nucleophile reacts with the more substituted carbon because it is this carbon that holds a greater degree of positive charge and electrostatics (carbocation stability) take a dominant role in determining the mechanism.

**EXAMPLE**

Predict the major product(s) of the ring opening reaction that occurs when the epoxide shown below is treated with:

a. ethanol and a small amount of sodium hydroxide  
   b. ethanol and a small amount of sulfuric acid

*Hint: be sure to consider both regiochemistry and stereochemistry!*

**Solution with an explanation**

a) [Diagram showing the reaction mechanism with S_N2 mechanism and back-side orientation of nucleophile and leaving group]  
   b) [Diagram showing the reaction mechanism with S_N1-S_N2 mechanism and nucleophile reacting at the more substituted carbon]
Addition of HX

Epoxides can also be opened by other anhydrous acids (HX) to form a trans halohydrin. When both the epoxide carbons are either primary or secondary the halogen anion will attack the less substituted carbon and an $S_N2$ like reaction. However, if one of the epoxide carbons is tertiary, the halogen anion will primarily attack the tertiary carbon in an $S_N1$ like reaction.

**EXAMPLE**

![Reaction Diagrams]

1. Reaction of an epoxide with HCl in ether.
2. Formation of a trans halohydrin.
Exercise

1. Given the following, predict the product assuming only the epoxide is affected. (Remember stereochemistry)

![Diagram of a molecule with an epoxide ring and an HBr reagent](image)

2. Predict the product of the following, similar to above but a different nucleophile is used and not in acidic conditions. (Remember stereochemistry)

![Diagram of a molecule with an epoxide ring and a nucleophile reaction](image)

3. Epoxides are often very useful reagents to use in synthesis when the desired product is a single stereoisomer. If the following alkene were reacted with an oxyacid to form an epoxide, would the result be a enantiomerically pure? If not, what would it be?

![Diagram of a molecule with an epoxide ring](image)

Answer

1. [Blank line]
Note that the stereochemistry has been inverted

2.

First, look at the symmetry of the alkene. There is a mirror plane, shown here.

Then, think about the mechanism of epoxidation with an oxyacid, take for example mCPBA. The mechanism is concerted, so the original cis stereochemistry is not changed. This leads to "two" epoxides.
However, these two mirror images are actually identical due to the mirror plane of the cis geometry. It is a meso compound, so the final result is a single stereoisomer, but not a single enantiomer.

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