The concept that chemical processes may not evolve "to completion" and instead tend to stop mid-way results in an "equilibrium" where reactants, products and potentially intermediates might be present. Identifying the concentrations of these populations requires rigorous definition of new reaction parameters (e.g. the equilibrium constant).

- **Dynamic Equilibria**

- **Le Chatelier's Principle**

- **Physical Equilibria**

- **Chemical Equilibria**
Heterogeneous Equilibria

- Solubility

- Acid-Base Equilibria