There are three aspects that make acid/base chemistry important reactions to emphasize:

1. They are common reactions since hydrogen atoms are ubiquitous in compounds
2. The bonds that hydrogen make to other atoms are often labeled with weaker bond dissociation energies so that hydrogen ions can be more easily transferred in reactions
3. The differences in Gibbs Energy ($\Delta G$) of the products and reactions are comparable to make many acid-base reaction incomplete (reversible).

Consequently, the same infrastructure (thermodynamics and kinetics) that apply to other reactions (e.g., redox or combustion) also apply to acid/base chemistry.
Ionization Constants

- Polyprotic Acids & Bases

- Buffers

- Acid/Base Titrations
Acid and Base Indicators