Four Factors in Affecting SN2 Reactions

1. The nature of the leaving group (SN2 Reactions-The Leaving Group)
2. The reactivity of the nucleophile (SN2 Reactions-The Nucleophile)
3. The solvent (SN2 Reactions-The Nucleophile)
4. The structure of the alkyl portion of the substrate (SN2 Reactions-The Substrate)

The Reactivity of the Nucleophile

Now that we have determined what will make a good leaving group, we will now consider nucleophilicity. That is, the relative strength of the nucleophile. Nucleophilicity depends on many factors, including charge, basicity, solvent, polarizability, and the nature of the substituents.

Increasing the Negative Charge Increases Nucleophilicity

Nucleophiles can be neutral or negatively charged. In either case, it is important that the nucleophile be a good Lewis base, meaning it has electrons it wants to share. The following diagram is just a reminder of some of the nucleophiles that were presented in the section covering nucleophilic substitution. In looking at these two types of nucleophiles, you should notice that a reactive atom, such as oxygen, in a neutral species can also be a reactive atom in a negatively charged species. For example, the O in OH is negatively charged, but the O in H2O is neutral.

**Neutral Nucleophiles**

\[ H_2O, \text{NH}_3, \text{RNH}_2, \text{R}_2\text{NH}, \text{R}_3\text{N}, \text{ROH}, \text{RCOOH}, \text{RSH}, \text{and PR}_3 \]

**Charged Nucleophiles**

\[ \text{RO}^-, \text{NH}_2^-, \text{RNH}^-, \text{R}_2\text{N}^-, \text{HS}^-, \text{RS}^-, \text{RSe}^-, \text{Cl}^-, \text{Br}^-, \text{I}^-, \text{F}^-, \text{CN}^-, \text{OH}^-, \text{RCO}_2^- \]

It has been experimentally shown that a nucleophile containing a negatively charged reactive atom is better than a nucleophile containing a reactive atom that is neutral. The next diagram illustrates this concept. Notice that when oxygen is part of the hydroxide ion, it bears a negative charge, and when it is part of a water molecule, it is neutral. The O of \text{OH} is a better nucleophile than the O of H2O, and results in a faster reaction rate. Similarly, when nitrogen is part of NH2, it bears a negative charge, and when it is part of NH3, it is neutral. The N of NH2 is a better nucleophile than the N of NH3, and results in a faster reaction rate.
When Moving Across a Row, Nucleophilicity Follows Basicity

To say that nucleophilicity follows basicity across a row means that, as basicity increases from right to left on the periodic table, nucleophilicity also increases. As basicity decreases from left to right on the periodic table, nucleophilicity also decreases. When it comes to nucleophilicity, do not assign this same rule when making comparisons between the halogens located in a column. In this case of moving up and down a column, nucleophilicity does not always follow basicity. It depends on the type of solvent you are using.

![Diagram](image)

**Basidity Increases**
**Nucleophilicity Increases**

In the section Nucleophilic Substitution, we assigned a relationship to leaving groups containing C, N, O, and F, showing that the strength of the leaving group follows electronegativity. This is based on the fact that the best leaving groups are those that are weak bases that do not want to share their electrons. The best nucleophiles however, are good bases that want to share their electrons with the electrophilic carbon. The relationship shown below, therefore, is the exact opposite of that shown for the strength of a leaving group.

- $\text{CH}_3^- > \text{NH}_2^- > \text{OH}^- > \text{F}^-$
  - Best
  - Worst

Solvents and Nucleophilicity

In general chemistry, we classified solvents as being either polar or nonpolar. Polar solvents can be further subdivided into protic and aprotic solvents.
Protic Solvents

A **protic solvent** is a solvent that has a hydrogen atom bound to an oxygen or nitrogen. A few examples of protic solvents include H₂O, ROH, RNH₂, and R₂NH, where water is an example of an inorganic protic solvent, and alcohols and amides are examples of organic solvents. The diagram below shows a few examples of protic solvents we will see.

![Diagram of polar protic solvents](image)

Since oxygen and nitrogen are highly electronegative atoms, the O-H and N-H bonds that are present in protic solvents result in a hydrogen that is positively polarized. When protic solvents are used in nucleophilic substitution reactions, the positively polarized hydrogen of the solvent molecule can interact with the negatively charged nucleophile. In solution, molecules or ions that are surrounded by these solvent molecules are said to be **solvated**. Solvation is the process of attraction and association of solvent molecules with ions of a solute. The solute, in this case, is a negatively charged nucleophile.

The following diagram depicts the interaction that can occur between a protic solvent and a negatively charged nucleophile. The interactions are called **hydrogen bonds**. A hydrogen bond results from a dipole-dipole force between between an electronegative atom, such as a halogen, and a hydrogen atom bonded to nitrogen, oxygen or fluorine. In the case below, we are using an alcohol (ROH) as an example of a protic solvent, but be aware that this interaction can occur with other solvents containing a positively polarized hydrogen atom, such as a molecule of water, or amides of the form RNH₂ and R₂NH.

![Diagram of hydrogen bonds](image)

**Why is this important?** Solvation weakens the nucleophile; that is, solvation decreases nucleophilicity. This is because the solvent forms a "shell" around the nucleophile, impeding the nucleophile's ability to attack an electrophilic carbon.
Furthermore, because the charge on smaller anions is more concentrated, **small anions are more tightly solvated than large anions**.

The picture below illustrates this concept. Notice how the smaller fluoride anion is represented as being more heavily solvated than the larger iodide anion. This means that the fluoride anion will be a weaker nucleophile than the iodide anion. In fact, it is important to note that fluoride will not function as a nucleophile at all in protic solvents. It is so small that solvation creates a situation whereby fluoride’s lone pair of electrons are no longer accessible, meaning it is unable to participate in a nucleophilic substitution reaction.

Previously we learned how nucleophilicity follows basicity when moving across a row. In our discussion on the effect of protic solvents on nucleophilicity, we learned that solvation weakens the nucleophile, having the greatest effect on smaller anions. In effect, when using protic solvents, nucleophilicity does not follow basicity when moving up and down a column. In fact, it's the exact opposite: when basicity **decreases**, nucleophilicity **increases** and when basicity **increases**, nucleophilicity **decreases**.

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**Aprotic Solvents**

An **aprotic solvent** is a solvent that lacks a positively polarized hydrogen. The next diagram illustrates several polar aprotic solvents that you should become familiar with.
Aprotic solvents, like protic solvents, are polar but, because they lack a positively polarized hydrogen, they do not form hydrogen bonds with the anionic nucleophile. The result, with respect to solvation, is a relatively weak interaction between the aprotic solvent and the nucleophile.

The consequence of this weakened interaction is two-fold. First, by using an aprotic solvent we can raise the reactivity of the nucleophile. This can sometimes have dramatic effects on the rate at which a nucleophilic substitution reaction can occur. For example, if we consider the reaction between bromoethane and potassium iodide, the reaction occurs 500 times faster in acetone than in methanol.

A second consequence that results from the weak interaction that occurs between aprotic solvents and nucleophiles is that, under some conditions, there can be an inversion of the reactivity order. An inversion would result in nucleophilicity following basicity up and down a column, as shown in the following diagram. When we considered the effects of protic solvents, remember that the iodide anion was the strongest nucleophile. Now, in considering aprotic solvents under some conditions, the fluoride anion is the strongest nucleophile.
Increasing Atomic Size Increases Nucleophilicity

Thus far, our discussion on nucleophilicity and solvent choice has been limited to negatively charged nucleophiles, such as F⁻, Cl⁻, Br⁻, and I⁻. With respect to these anions we learned that, when using protic solvents, nucleophilicity does not follow basicity, and when using aprotic solvents, the same relationship can occur, or there could be an inversion in the order of reactivity.

What happens as we move up and down a column when considering uncharged nucleophiles? It turns out that, in the case of uncharged nucleophiles, size dictates nucleophilicity. This is because larger elements have bigger, more diffuse, and more polarizable electron clouds. This cloud facilitates the formation of a more effective orbital overlap in the transition state of bimolecular nucleophilic substitution (SN₂) reactions, resulting in a transition state that is lower in energy and a nucleophilic substitution that occurs at a faster rate.

![Size Increases, Basicity Decreases, Nucleophilicity Increases](image)

Examples of Uncharged Nucleophiles

- \( \text{H}_2\text{Se} > \text{H}_2\text{S} > \text{H}_2\text{O} \)
- \( \text{PH}_3 > \text{NH}_3 \)

Sterically Hindered Nucleophiles React More Slowly

In the section Kinetics of Nucleophilic Substitution Reactions, we learned that the SN₂ transition state is very crowded. Recall that there are a total of 5 groups around the electrophilic center.
For this reason, stericly hindered nucleophiles react more slowly than those lacking steric bulk.

Next section: SN$_2$ Reactions-The Substrate

Contributors

- Rachael Curtis (UCD)