



Part A

Solution	Relevant Observation 1	Relevant observation 2	pH est

Could you have predicted the pH of some of these solutions? Explain.

Part B

Solution	pH	Equation showing the substance reacting as acid or base

Part C

Quantity	Value for 0.1 M acetic acid	Value for 0.1 M acetate
pH		
[H ₃ O ⁺] at equilibrium in mol/L		
[OH ⁻] at equilibrium in mol/L		
[acetic acid] at equilibrium in mol/L		
[acetate] at equilibrium in mol/L		
K		
pKa		

Part D

What pH are you trying to set the buffer to: pH = _____

What volumes of acetic acid and acetate will you use?

Show the pH calculation

What pH did you measure?

Reflect on your experience today (e.g. which calculation was the most challenging for you)