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| Group Name: |  |

VSEPR

Group Lab Report

## **LibreTexts page:** [**7: Molecular Structure**](https://chem.libretexts.org/Courses/University_of_Arkansas_Little_Rock/Chem_1402%3A_General_Chemistry_1_%28Belford%29/Laboratory/07%3A_Experiment_7_-_Molecular_Structure)

## **(**<https://chem.libretexts.org/link?214684>)

**Please don’t edit, rearrange or delete anything that is already in this document. Just add your answers inside the boxes.**

**You can use shortcuts for superscripts and subscripts when needed:**

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**Use the VSEPR Worksheet to answer the questions.**

**To indicate the dipole, ROTATE the image to position atoms with partial positive charge on top.**

**Use the first compound as example.**

|  |  |  |
| --- | --- | --- |
| **Compound** | **Name electronic geometry and identify hybrid orbitals** | **Name molecular geometry, indicate bond angle and if polar. Paste the image and orient it so partial positive charge is at the top.**  |
| **H2O** | Tetrahedral, sp3 | Bent, 104.5 deg, polar |
| **NH3** |  |  |
| **NH4+** |  |  |
| **SO2** |  |  |
| **NO3-** |  |  |
| **SeCl4** |  |  |
| **AlCl3** |  |  |
| **XeCl4** |  |  |